



# Cambridge IGCSE™

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## CHEMISTRY

0620/31

Paper 3 Theory (Core)

May/June 2024

1 hour 15 minutes

You must answer on the question paper.

No additional materials are needed.

### INSTRUCTIONS

- Answer **all** questions.
- Use a black or dark blue pen. You may use an HB pencil for any diagrams or graphs.
- Write your name, centre number and candidate number in the boxes at the top of the page.
- Write your answer to each question in the space provided.
- Do **not** use an erasable pen or correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.
- You should show all your working and use appropriate units.

### INFORMATION

- The total mark for this paper is 80.
- The number of marks for each question or part question is shown in brackets [ ].
- The Periodic Table is printed in the question paper.

This document has **20** pages. Any blank pages are indicated.

1 Fig. 1.1 shows the structures of seven substances, **A**, **B**, **C**, **D**, **E**, **F** and **G**.

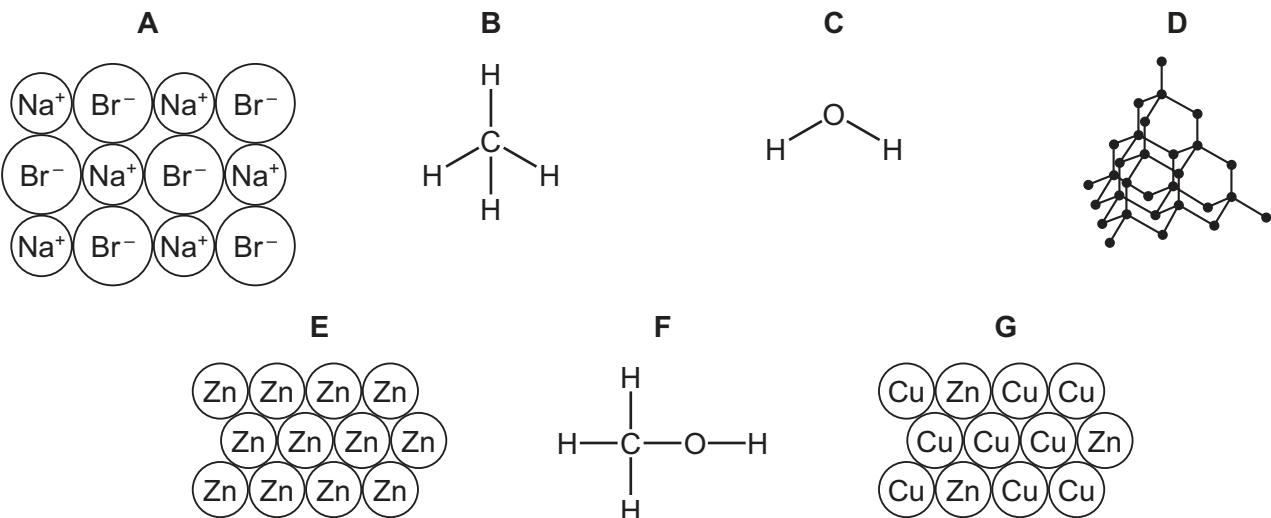


Fig. 1.1

(a) Answer the following questions using only the structures in Fig. 1.1.  
Each structure may be used once, more than once or not at all.

State which structure represents:

(i) an alloy

..... [1]

(ii) a substance that only conducts electricity when molten or in aqueous solution

..... [1]

(iii) a giant covalent structure

..... [1]

(iv) a compound that is a product formed in a hydrogen–oxygen fuel cell

..... [1]

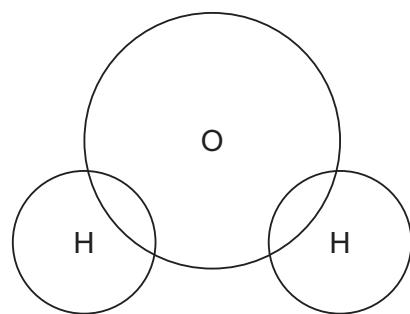
(v) a compound with a high melting point

..... [1]

(vi) a gas that is responsible for increased global warming.

..... [1]

(b) Complete Fig. 1.2 to show the dot-and-cross diagram for structure C.  
Show the outer shell electrons only.



**Fig. 1.2**

[2]

[Total: 8]

2 (a) Table 2.1 shows the percentages by mass of the elements present in the human body.

**Table 2.1**

element	percentage by mass of element
calcium	1.50
carbon	18.00
chlorine	0.15
hydrogen	10.00
magnesium	0.05
nitrogen	3.00
oxygen	65.00
phosphorus	1.00
potassium	0.35
sodium	0.15
sulfur	0.25
other elements	0.55

Answer these questions using information from Table 2.1.

(i) Name the non-metallic element in Table 2.1 that has the lowest percentage by mass.

..... [1]

(ii) Name an element in Table 2.1 that is in Period 4 of the Periodic Table.

..... [1]

(b) Some medicines contain a compound made of  $Mg^{2+}$  ions and  $OH^-$  ions.

Name the compound made of  $Mg^{2+}$  ions and  $OH^-$  ions.

..... [1]

(c) Describe the observations when aqueous sodium hydroxide is added dropwise to a solution containing calcium ions until the sodium hydroxide is in excess.

observations with dropwise addition of sodium hydroxide

.....  
observations with excess sodium hydroxide

..... [2]

(d) Name a calcium salt that is soluble in water.

..... [1]

(e) Table 2.2 shows some properties of the Group I metals.

**Table 2.2**

metal	melting point /°C	observations on reaction with water
lithium	181	
sodium	98	bubbles form rapidly but no flame
potassium		bubbles form very rapidly and flame
rubidium	39	explodes

Use the information in Table 2.2 to predict:

- the melting point of potassium

.....

- the observations when lithium reacts with water.

.....

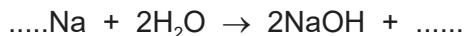
[2]

(f) State how the density of the Group I elements changes down the group.

..... [1]

(g) Sodium reacts with water to produce sodium hydroxide and a gas which pops with a lighted splint.

Complete the symbol equation for this reaction.



[2]

[Total: 11]

3 Aluminium is extracted by electrolysis of its purified ore.

(a) Name the main ore of aluminium.

..... [1]

(b) Fig. 3.1 shows the apparatus used in the extraction of aluminium.

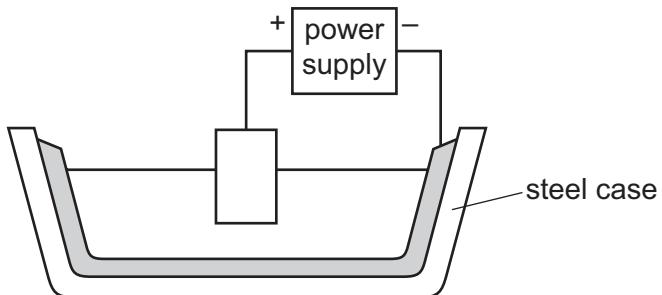


Fig. 3.1

(i) Label the cathode in Fig. 3.1.

[1]

(ii) The electrolyte contains molten aluminium oxide.

State the product formed at each electrode.

positive electrode .....

negative electrode .....

[2]

(c) State **two** physical properties that explain why aluminium is used in overhead electrical cables.

1 .....

2 .....

[2]

(d) Aluminium ore is purified by reacting it with sodium hydroxide.  
Sodium hydroxide is an alkali.

(i) State the meaning of the term alkali.

..... [1]

(ii) Describe how to find the pH of a dilute solution of sodium hydroxide using universal indicator paper.

.....

..... [2]

(iii) A dilute solution of sodium hydroxide is added to a solution of methyl orange in acid until the sodium hydroxide is in excess.

State the colour change of the methyl orange.

from ..... to ..... [2]

[Total: 11]

4 (a) Fig. 4.1 shows the displayed formula of a compound extracted from a plant.

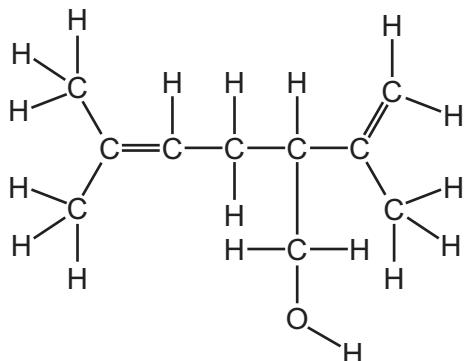


Fig. 4.1

On Fig. 4.1, draw a circle around **one** functional group that makes this compound unsaturated. [1]

(b) A student extracts mixtures of coloured compounds from four different plants, **Q**, **R**, **S** and **T**.

Fig. 4.2 shows the results of chromatography of these mixtures using an organic solvent.

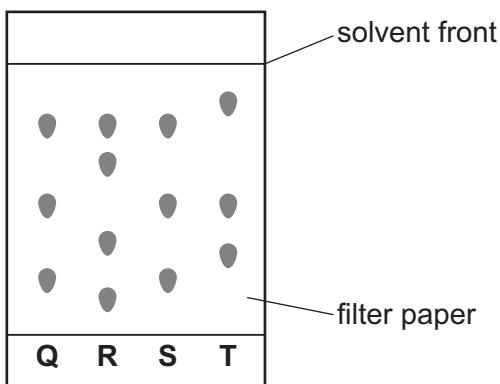


Fig. 4.2

(i) Deduce which plant, Q, R, S or T, contains the greatest number of coloured compounds.

---

[11]

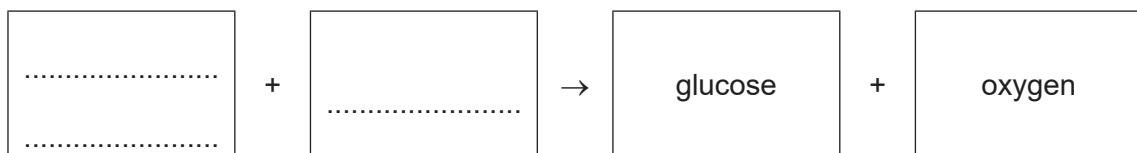
(ii) Deduce which **two** plants, Q, R, S or T, contain exactly the same coloured compounds.

..... and ..... [11]

(iii) State the meaning of the term solvent.

(c) (i) Plants produce glucose and oxygen by photosynthesis.

Complete the word equation for photosynthesis.



[2]

(ii) Name one **other** substance that is essential for photosynthesis.

..... [1]

[Total: 7]

5 (a) An atom of carbon is represented by the symbol shown.



Describe this atom of carbon in terms of:

- the position of the electrons, neutrons and protons in this atom

.....  
.....  
.....

- the number of neutrons and number of protons

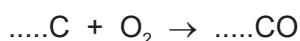
.....  
.....  
.....

- the electronic configuration.

.....

[5]

(b) (i) Complete the symbol equation for the incomplete combustion of carbon to produce carbon monoxide.



[2]

(ii) State **one** adverse effect of carbon monoxide.

..... [1]

(c) Fig. 5.1 shows the displayed formula of chromium carbonyl.

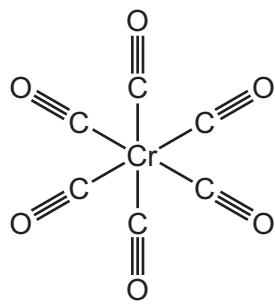


Fig. 5.1

Deduce the molecular formula of chromium carbonyl.

..... [1]

(d) Another compound of chromium has the formula  $\text{Na}_2\text{Cr}_2\text{C}_{10}\text{O}_{10}$ .

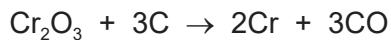
Complete Table 5.1 to calculate the relative molecular mass of  $\text{Na}_2\text{Cr}_2\text{C}_{10}\text{O}_{10}$ .

**Table 5.1**

type of atom	number of atoms	relative atomic mass	
sodium	2	23	$2 \times 23 = 46$
chromium		52	
carbon		12	
oxygen		16	

$$\text{relative molecular mass} = \dots \quad [2]$$

(e) Chromium can be produced by heating chromium(III) oxide,  $\text{Cr}_2\text{O}_3$ , with carbon.



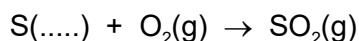
Describe how this equation shows that chromium(III) oxide is reduced.

..... [1]

[Total: 12]

6 Large pieces of solid sulfur burn in excess oxygen to produce sulfur dioxide gas.

(a) Complete the equation by adding the missing state symbol.



[1]

(b) Fig. 6.1 shows how the mass of sulfur changes as the reaction proceeds.

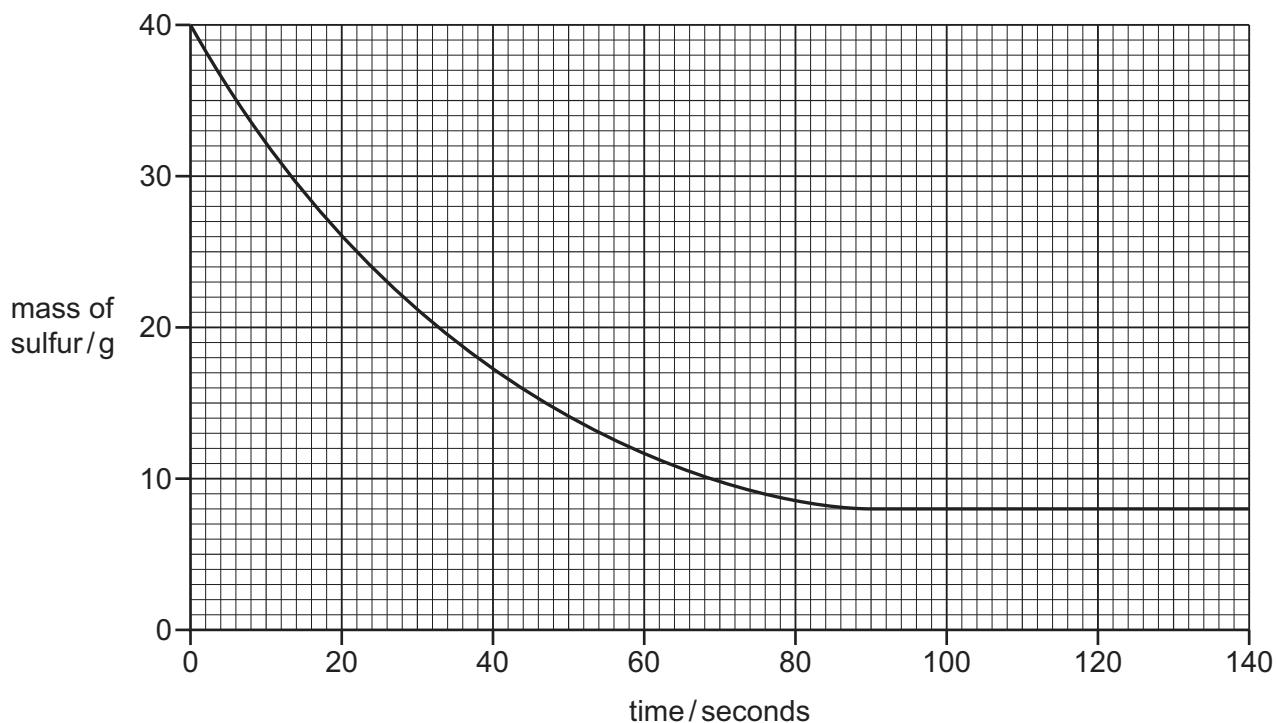


Fig. 6.1

Deduce the time taken for the reaction to finish.

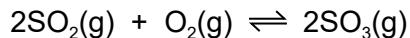
..... [1]

(c) The experiment is repeated using powdered sulfur.

Describe the effect on the rate of reaction of using powdered sulfur rather than large pieces of sulfur.

..... [1]

(d) Sulfur dioxide reacts with oxygen in a closed container.



(i) Describe the effect, if any, each of the following has on the rate of this reaction.

All other conditions stay the same.

- The temperature is decreased.

.....

- The pressure of the gases is increased.

.....

[2]

(ii) Changing concentration changes the rate of a reaction.

Choose the correct unit of concentration from the list.

Draw a circle around your chosen answer.

**dm<sup>3</sup>/mol**

**mol/dm**

**mol/dm<sup>2</sup>**

**mol/dm<sup>3</sup>**

[1]

(e) Sulfur dioxide is an air pollutant.

(i) State **one** adverse effect of sulfur dioxide.

.....

[1]

(ii) Emissions of sulfur dioxide can be reduced by using low-sulfur fossil fuels.

State one **other** way of reducing sulfur dioxide emissions from fossil fuels.

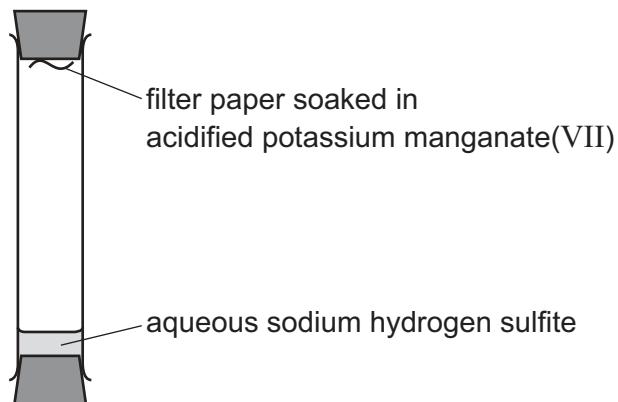
.....

[1]

(f) Aqueous sodium hydrogen sulfite releases sulfur dioxide gas at room temperature.

Sulfur dioxide changes the colour of acidified potassium manganate(VII) from purple to colourless.

Fig. 6.2 shows a sealed tube with a small volume of aqueous sodium hydrogen sulfite at the bottom. A piece of filter paper soaked in acidified potassium manganate(VII) is attached to the top of the tube.



**Fig. 6.2**

The filter paper remains purple at first.

The filter paper becomes colourless after a short time.

Explain these results in terms of kinetic particle theory.

.....

.....

.....

.....

[3]

[Total: 11]

7 Iron is a metal. Iron has a high density, a high melting point and a high boiling point.

(a) State three **other** physical properties of iron.

1 .....

2 .....

3 .....

[3]

(b) (i) State the conditions needed for iron to rust.

..... [2]

(ii) Rust is hydrated iron(III) oxide.

State if iron(III) oxide is an acidic or basic oxide.

Give a reason for your answer.

..... [1]

(iii) Complete this sentence about methods of preventing rusting.

Rusting can be prevented by painting or .....

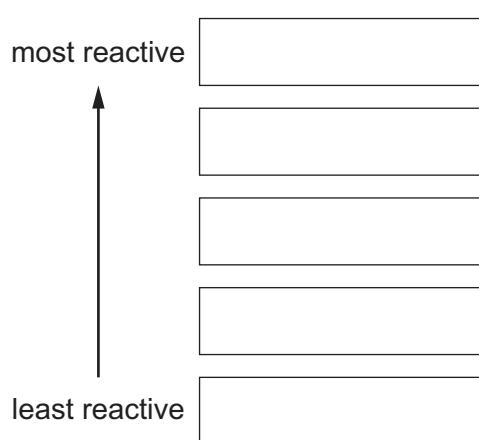
..... [1]

(c) The list shows five metals.

calcium	copper	iron	silver	sodium
---------	--------	------	--------	--------

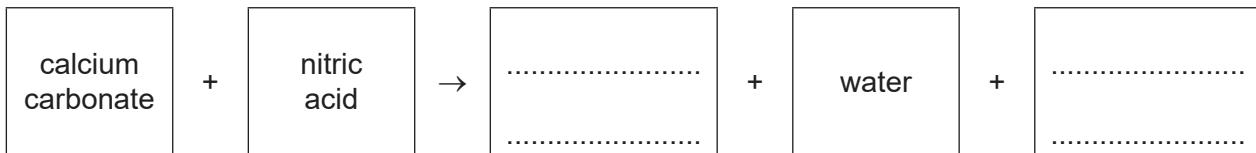
Put these metals in order of their reactivity.

Put the most reactive metal at the top.



[2]

(d) Complete the word equation for the reaction of calcium carbonate with nitric acid.



[2]

[Total: 11]

8 (a) Fig. 8.1 shows the displayed formulae of five organic compounds, **V**, **W**, **X**, **Y** and **Z**.

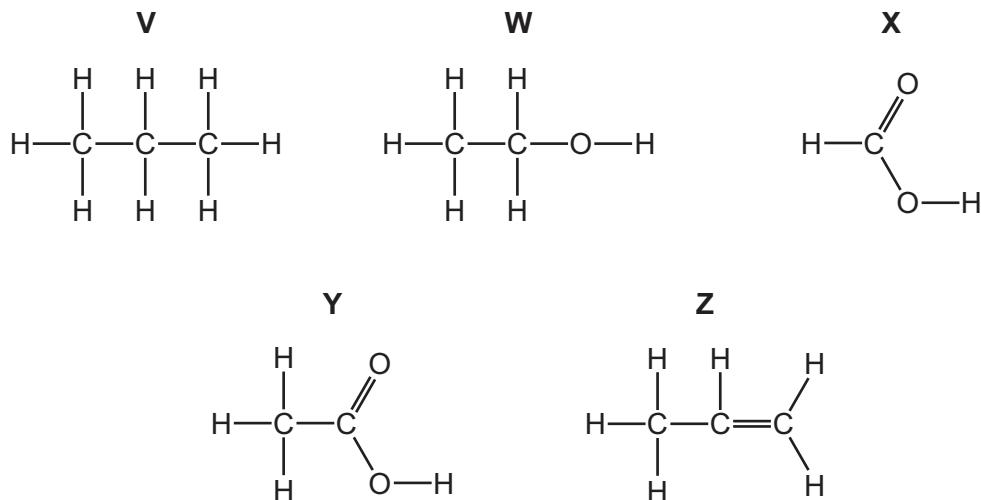


Fig. 8.1

(i) State which **two** of the compounds, **V**, **W**, **X**, **Y** and **Z**, are in the same homologous series.

..... and ..... [1]

(ii) Explain why compound **V** is an alkane.

.....  
..... [2]

(iii) State the name of the homologous series to which compound **X** belongs.

..... [1]

(b) Ethanol can be manufactured by the catalytic addition of steam to ethene.

(i) State the temperature and pressure required for this reaction.

temperature ..... °C

pressure ..... atm

[2]

(ii) Name one **other** method of manufacturing ethanol.

..... [1]

(c) Describe how alkenes are manufactured from petroleum fractions.

.....  
.....  
..... [2]

[Total: 9]



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# The Periodic Table of Elements

I		II		Group																						
				I						II			III			IV		V		VI		VII		VIII		
				Key																						
3	4	Li	Be	atomic number	1	H	hydrogen																			
lithium	beryllium	7	9	name																						
7	9	11	12	relative atomic mass																						
Na	Mg	magnesium	23																							
19	20	21	22																							
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Zn	Ga															
potassium	calcium	scandium	titanium	vanadium	chromium	manganese	iron	cobalt	nickel	zinc	gallium															
39	40	45	48	51	52	55	56	59	59	64	70															
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd															
85	88	89	91	93	96	—	101	103	106	108	112	indium														
57	38	39	40	41	42	43	44	45	46	47	48	49	50	51	52	53	54	54	54	54	54	54	54	54		
La	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Zn	Ga																
139	58	21	22	23	24	25	26	27	28	29	30	31	32	33	34	35	36	36	36	36	36	36	36	36		
lanthanum	Ce	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Zn	Ga															
140	59	60	61	62	63	64	65	66	67	68	69	70	71	72	73	74	75	76	77	78	79	80	81	82		
Pr	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu													
141	61	60	61	62	63	64	65	66	67	68	69	70	71	72	73	74	75	76	77	78	79	80	81	82		
Th	Pa	U	Pa	U	U	Am	Cm	Bk	Cf	Fm	Md	No	Lu													
232	91	92	93	94	95	96	97	98	99	100	101	102	103													
Ac	Th	Pa	U	U	U	U	U	U	U	U	U	U	U	U	U	U	U	U	U	U	U	U	U	U		
—	231	238	231	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	

57	58	59	60	61	62	63	64	65	66	67	68	69	70	71	72	73	74	75	76	77	78	79	80	81	82	
La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu												
139	140	141	140	141	144	144	144	145	143	145	147	147	147	147	147	147	147	147	147	147	147	147	147	147	147	
Praseodymium	Cerium	Neodymium	Pr	Neodymium	Pr																					
89	90	91	92	93	94	95	96	97	98	99	100	101	102	103												
Ac	Th	Pa	U	Pa	U	U	U	U	U	U	U	U	U	U	U	U	U	U	U	U	U	U	U	U	U	
—	232	231	238	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	—	

The volume of one mole of any gas is 24 dm<sup>3</sup> at room temperature and pressure (r.t.p.).